Sanjay Ghodawat University, Kolhapur Established as State Private University under Govt. of Maharashtra. Act No XL, 2017		2019-20 EXM/P/09/00
Year and Program: F Y B. Pharm	**School of Pharmaceutical Sciences**	**Department : Pharmacy**
Course Code : BP102T	**Course Title :Pharmaceutical Analysis - I**	**Semester – I**
Date: Tuesday **Day : 23.02.2021**	**Examination: End Semester**	**Time: 2 Hrs** **Max Marks: 50**

Instructions: 1) All questions are compulsory

2) Draw neat labeled diagram if necessary.

			Marks	Bloom's Level	CO
Q.1		**Objective Type Questions (Any Seven)**	**14**		
	a)	Define Normality, Molarity with their formula.	02	1	1
	b)	What is error? And Enlist there types.	02	1	1
	c)	What is acid and base? According to Lewis and Bronsted -Lowry theory.	02	1	1
	d)	What is non- aqueous titration? And write advantages and disadvantages of it.	02	1	1
	e)	Classify the ligands incomplexometric titration	02	4	1
	f)	What is principle of gravimetric analysis? Write factors affecting precipitation in gravimetric analysis.	02	1	2
	g)	Define Redox titration and enlist two name of indicator	02	1	1
	h)	What is conductance and specific resistance?	02	1	1
	i)	What is iodometry?	02	1	1
	j)	Write the applications of potentiometric analysis.	02	1	3
Q.2		**Long Answers (Answer 2 out of 3)**	**15**		
	a)	Illustrate theories of acid and base, Types and theories of indicators in acid base titration.	7.5	2	2
	b)	Describe briefly about iodometric and iodimetric titration. Give example.	7.5	2	2
	c)	Discuss in brief Fajan's method and Differentiate between Mohr`s method and Volhard`s method.	7.5	2	2
Q3		**Short Answers (Answer 7 out of 9)**	**21**		
	a)	What is primary standard and secondary standard? Give their characteristics and example.	03	2	1
	b)	Explain in brief methods for minimizing errors.	03	2	1
	c)	Define limit test?Write limit test for chloride.	03	1	1
	d)	Draw neat labeled diagram of Saturated Calomelelectrode (SCE) and write application for	03	3	4

	Potentiometric titration.			
e)	Describe the principle and application of polarography.	03	2	3
f)	Write types of redox titration and discuss any two methods.	03	2	1
g)	Describe impurities in precipitates and write assay of Barium sulphate.	03	2	2
h)	Explain in brief solvents used in non-aqueous titration.	03	2	1
i)	Draw well-labeled assembly diagram and enlist the different electrode in potentiometric titration.	03	3	4